

Student Roll Number

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Example Student Roll No.

1	3	5	2	4	6
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0	0	0	0	0	0
1	1	1	1	1	1
2	2	2	2	2	2
3	3	3	3	3	3
4	4	4	4	4	4
5	5	5	5	5	5
6	6	6	6	6	6
7	7	7	7	7	7
8	8	8	8	8	8
9	9	9	9	9	9

0	0	0	0	0	0
●	1	1	1	1	1
2	2	2	●	2	2
3	●	3	3	3	3
4	4	4	4	●	4
5	5	●	5	5	5
6	6	6	6	6	●
7	7	7	7	7	7
8	8	8	8	8	8
9	9	9	9	9	9

Sign. and Seal of Supdtt.

Paper: CHEMISTRY

Part: 9th

Time: 20 Minutes

Marks: 12

Exam Code: 9171

NOTE
FILL IN THE
CORRECT CIRCLE ONLY

- A Cation is:
A) natural, B) negatively charged, C) positively charged, D) coloured,
- The noble metals are very:
A) reactive, B) passive, C) soft, D) electronegative,
- Ionization potential of hard metals is:
A) low, B) high, C) zero, D) negative,
- 10 gm of solute present in 100 gm of solution is called:
A) v/w %, B) v/v %, C) w/v %, D) w/w %,
- Water droplet in air is an example of solution of:
A) gas in gas, B) gas in liquid, C) liquid in gas, D) liquid in liquid,
- Which one of the following is an electrolyte?
A) sugar solution, B) table salt solution, C) plastic, D) wood,
- The number of valence electrons of Group-II elements is:
A) 2, B) 4, C) 6, D) 8,
- Which kind of bond exists in HCl?
A) purely ionic, B) polar covalent, C) covalent, D) coordinate covalent,
- Which metal occurs in liquid state?
A) Na, B) Ca, C) Hg, D) Pt,
- The chemistry which deals with the qualitative and quantitative analysis of matter is called chemistry. A) organic, B) inorganic, C) analytical, D) industrial,
- The number of molecules in 0.5 mole of glucose is:
A) 3.01×10^{23} , B) 12.04×10^{23} , C) 3.01×10^{-23} , D) none of these,
- The chemical properties of an atom depend upon: A) no. of neutrons in nucleus, B) no. of protons in nucleus, C) no. of electrons in outermost shell, D) all these,

A	B	●	D
A	●	C	D
●	B	C	D
A	B	C	●
A	B	●	D
A	●	C	D
●	B	C	D
A	●	C	D
A	B	●	D
A	B	●	D
A	B	C	●
A	B	●	D

5171
CHEMISTRY — 9th

Time: 2 Hours 40 Minutes

SECTION-B

Marks: 32

1. Attempt any eight of the following. All carry equal marks.
- i. Calculate the molecular mass of sugar.
 - ii. What do you mean by atomic number? Give example.
 - iii. Give electronic configuration of ^{10}A , ^{12}B .
 - iv. What are the isotopes of chlorine? Draw their structure.
 - v. Explain duplet and octet rule with examples.
 - vi. Write physical properties of solid.
 - vii. Define solubility and molarity.
 - viii. Find the oxidation number of sulphur in H_2SO_4 and nitrogen in HNO_3 .
 - ix. What are weak and strong electrolytes? Give two examples of each.
 - x. Differentiate between metals and non-metals.
 - xi. Write physical properties of sodium.

SECTION-C

Marks: 21

NOTE: Attempt any three of the following questions. All questions carry equal marks.

2.
 - i. Give reasons that alloys are mixtures not compounds.
 - ii. What did Rutherford deduce from his experiments? Give diagram.
3.
 - i. What is electronegativity? Write its trend in periodic table.
 - ii. If 3dm^3 of air is heated from 300 K to 400 K at constant pressure then what is the volume of the gas at high temperature?
4.
 - i. How diffusion differs from effusion.
 - ii. Calculate the molarity of solution which has 0.530g of Na_2CO_3 in 250cm^3 of solution.
5.
 - i. Describe Daniell cell.
 - ii. Write down chemical properties of halogens.