

Serial No. Of the Answer Book
Roll No. (in figure)

(in Word)

Superintendent Seal & Signature

FIC. No (For office use only)

901601

CHEMISTRY (Fresh) 9th

FIC. No (For office use only)

Total Time: 3 Hours

Max: Marks: 65

Note: There are THREE Sections of this Paper i.e. A, B and C, attempt each according to the given instructions.

Time: 20 Minutes

SECTION-A

Marks: 12

Note: Attempt all parts of Section -- A. Section --A must be return to the superintendent after 20 minutes even if you have not attempted any question. Overwriting/ defacing/Cutting etc is prohibited in Section-A and no credit will be given to such answer.

I. Write the correct option i.e. A/B/C/D in the empty boxes.

- i. Which one of the following has both empirical and molecular formula _____ C
(A) C_6H_6 (B) H_2O_2 (C) H_2O (D) $C_6H_{12}O_6$
- ii. Which one is a homogeneous mixture _____ B
(A) Smoke (B) air (C) fog (D) smog
- iii. The no. of groups in periodic table are _____ C
(A) six (B) seven (C) eight (D) nine
- iv. Which pair of compounds ionizes completely in water _____ A
(A) NaCl and HCl (B) NaBr and CO_2
(C) NH_3 and HCl (D) CH_3COOH and K_2CO_3
- v. Which one of the following is formed by three types of bonds. _____ D
(A) H_2O (B) CO_2 (C) NaCl (D) NH_4Cl
- vi. The increase in temperature of the gases decrease the _____ C
(A) Pressure (B) Volume (C) Forces of attraction (D) Kinetic energy
- vii. Milk is an example of _____ B
(A) Solution (B) colloidal solution (C) suspension (D) Heterogenous mixture
- viii. The oxidation number of carbon in CO_2 is _____ C
(A) +2 (B) +3 (C) +4 (D) +6
- ix. Which one of the following is a strong electrolyte in solution _____ C
(A) H_2CO_3 (B) NH_4OH (C) HCl (D) CH_3COOH
- x. Sodium belongs to group _____ A
(A) I (B) II (C) III (D) IV
- xi. Which one is powerful reducing agent _____ B
(A) Li (B) Na (C) Mg (D) Ca
- xii. Hard metals possess ionization potential _____ B
(A) Low (B) High (C) Zero (D) Negative

Note: Time allowed for section B and C is 2 hours and 40 minutes.

SECTION "B"

Marks: 32

II. Attempt any EIGHT Parts out of the following. Each Part carries equal marks.

- i. What is meant by Atomic number?
- ii. How many grams are there in 3.5 moles of water H₂O ?
- iii. Why alloys are mixtures but not compounds?
- iv. Write Electronic configuration of Mg and p
- v. Write uses of isotopes.
- vi. Discuss the Electronegativity .
- vii. Write 4 properties of metals.
- viii. What is metallic bond?
- ix. State and explain Boyle's law.
- x. What is meant by solubility?
- xi. Discuss purposes of Electroplating

SECTION "C"

Marks: 21

Note: Attempt any THREE questions of the following. Each question carries equal Marks.

- III. (a) Explain the assumptions of Neil Bohr's Atomic theory. 4
(b) Define energy level and energy sub-level. 3
- IV. (a) Define covalent bond and explain its types in detail. 4
(b) Ionic compounds are good conductor as compared to the Covalant compounds why? 3
- (a) What is the difference between Oxidizing and reducing agents? Also explain with examples. 4
(b) Dry cell is an irreversible cell. Why? 3
- VI. (a) Write the chemical properties of sodium. 4
(b) Cu is used for making wires and sheets. Why? 3